

INTRODUCTION

- Nickel with oxygen forms three binary phases: NiO, NiO₂ and Ni₂O₃.
- The only known and well scientifically studied nickel oxide phase is NiO (green nickel) and the other two phases are weakly characterized.
- In 1961, Aggarwal identified the crystal structure of Ni₂O₃ as a hexagonal structure [1].
- There are no mentions in literature regarding the Ni₂O₅ oxide - this phase is completely unexplored.
- We focus on **theoretical prediction** of crystal and electronic structure and dynamic stability of Ni₂O₃ and Ni₂O₅ using **Density Functional Theory modelling (DFT)** and **Evolutionary algorithms (EA)**.

METHODS

- Spin-polarized DFT calculation were performed in program VASP [2] with GGA PBEsol functional and plane-wave cut-off energy 520 eV.
- Strong electron correlations were accounted for with LSDA+U method [3] with Hubbard parameter U_{Ni} = 5 eV in combination with Hund's correction J_{Ni} = 1 eV.
- Lattice dynamics was calculated for supercells using direct phonon method and quasiharmonic approximation as implemented in the program PHONOPY. [5].
- EA calculations were performed in XTalOpt program [6] in combination with DFT calculations (with PBE functional and cut-off 530 eV) and 1000 structures was generated for Z=2.
- Visualization of crystal structures was done in VESTA [4].

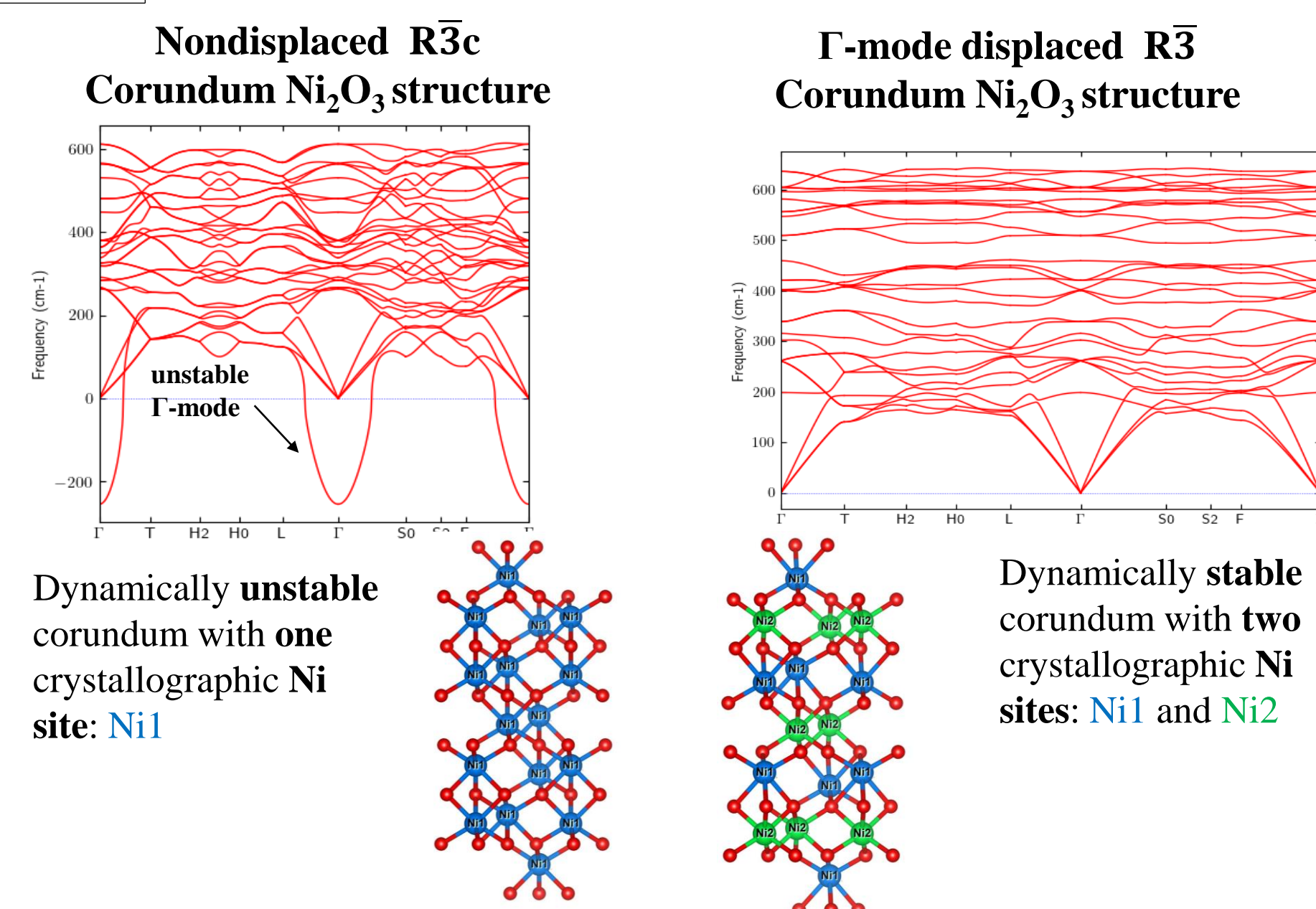
CONCLUSIONS

- Considering the 7 structure types commonly taken by transition metal sesquioxides, corundum structure (R3c) is favoured yet dynamically unstable.
- Dynamically stable Ni₂O₃ ground-state structure is predicted to be a non-centrosymmetric corundum type (R3, 148) with two crystallographic Ni1 and Ni2 sites.
- Ni₂O₃ ground state structure is predicted to be ferromagnetic semiconductor with band gap equal to 1.4 eV.
- EA calculations confirm the corundum type ground state Ni₂O₃ structure.

Ni₂O₅

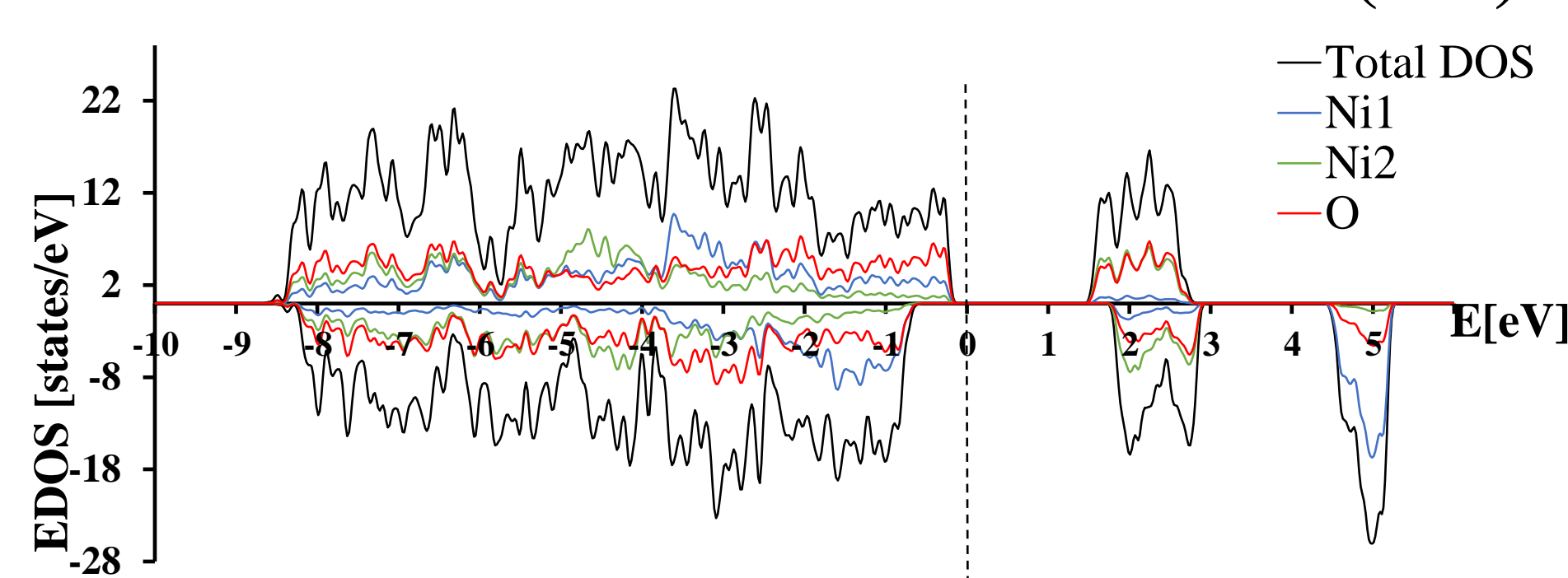
- Monoclinic structure (C2/c, 15) common to V, Nb and Ta pentoxides is favoured and predicted to be dynamically stable.
- O-O bridges are present in the monoclinic Ni₂O₅ structure contrary to the known V, Nb and Ta pentoxides.
- Ni₂O₅ in the C2/c structure is predicted to be ferromagnetic conductor.
- EA calculations further confirm the strong tendency for formation of O-O bonds in Ni₂O₅.
- Presence of O-O bonds in all Ni₂O₅ models suggest that it is unlikely to stabilize Ni⁵⁺ cation at ambient conditions in the Ni₂O₅ oxide.

Ni₂O₃ LATTICE DYNAMICS



- The corundum structure is dynamically unstable with respect to phonon soft mode at Γ point of the Brillouin zone.
- This vibrational mode lowers the symmetry of the corundum structure from R3c (167) to R3 (148) with two distinct crystallographic Ni sites.

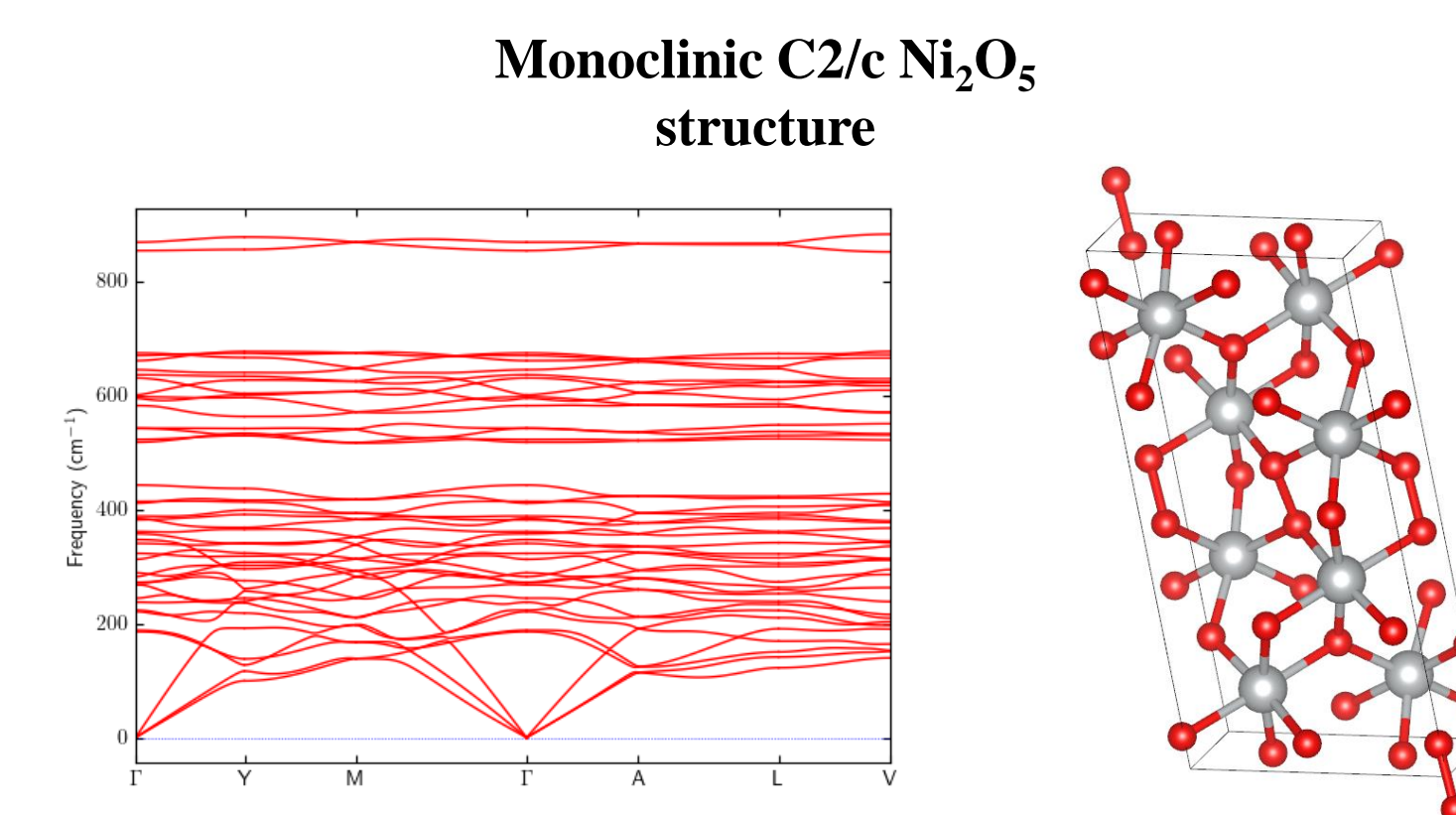
ELECTRONIC STRUCTURE (R3)



- Ferromagnetic semiconductor.
- Insulating band gap is 1.4 eV.
- R3c corundum structure** - two magnetically slightly distinct nickel sites, which are organized into separated layers in the same fashion as Ni1 and Ni2 in the distorted R3 corundum structure.
- R3 corundum structure** - magnetic moment on Ni1 is increased relative to the R3c corundum and magnetic moment on Ni2, which is characteristic of the very short Ni-O distances is severely reduced.

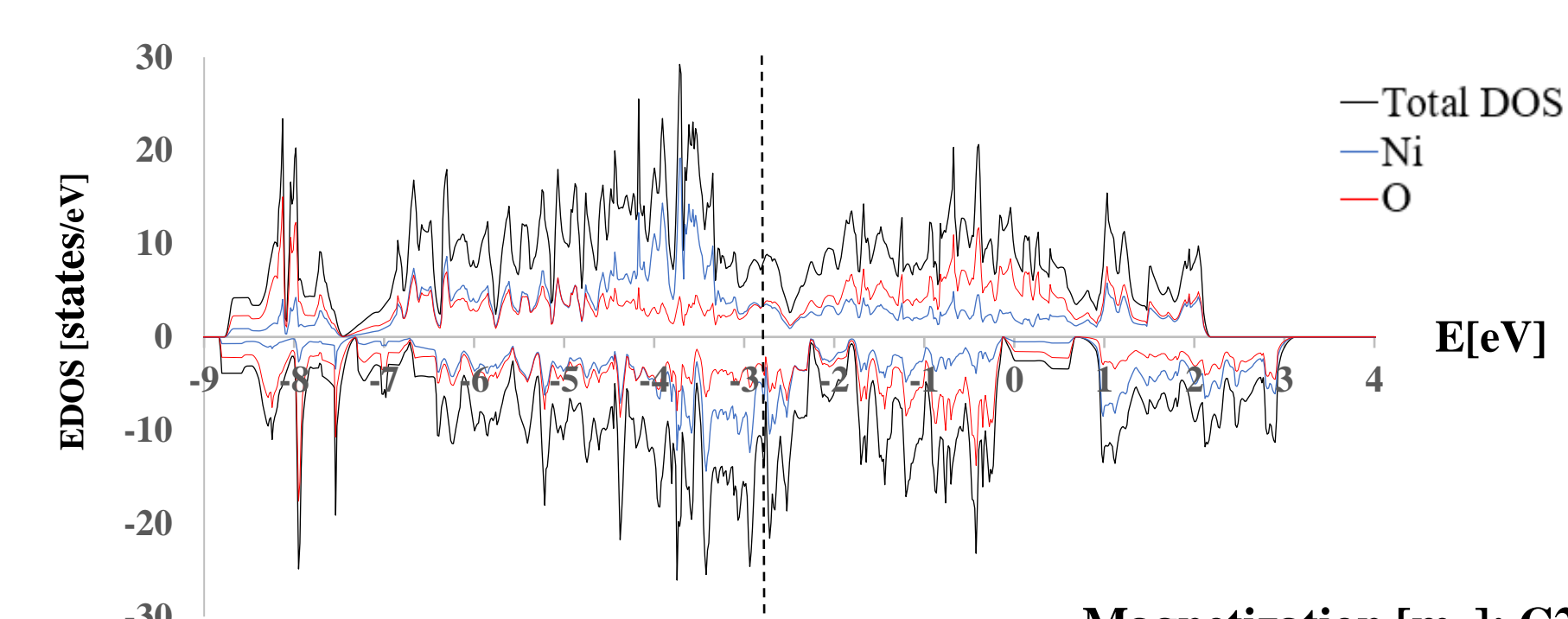
	Ni1	Ni2	O
nondisplaced R3c	1.138, 1.141	-	-0.072, -0.086, -0.099
displaced R3	1.575	0.361	0.013, 0.016, 0.025

Ni₂O₅ LATTICE DYNAMICS



- The monoclinic (C2/c, 15) structure is dynamically stable, there are no negative frequencies present.

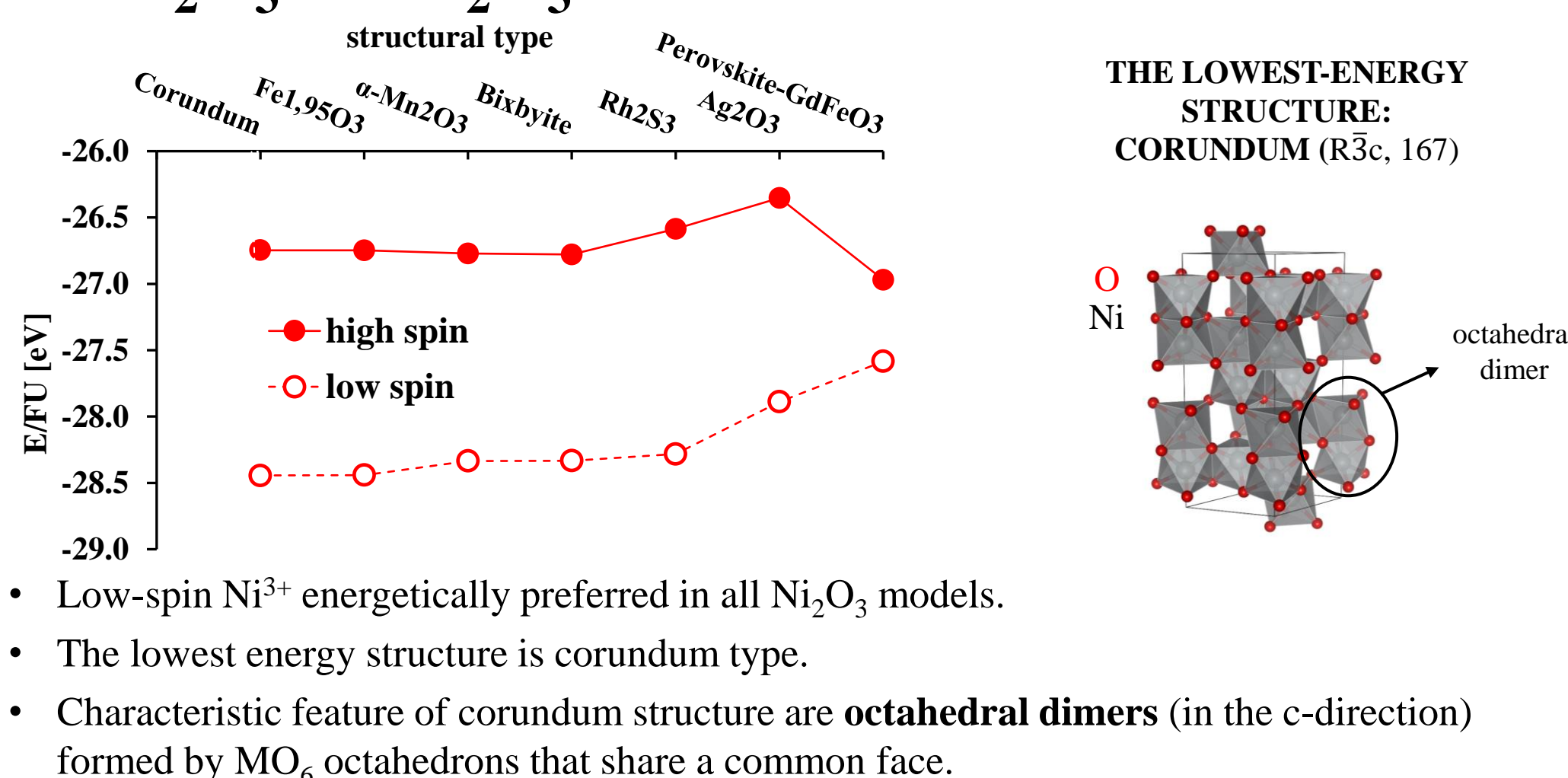
ELECTRONIC STRUCTURE (C2/c)



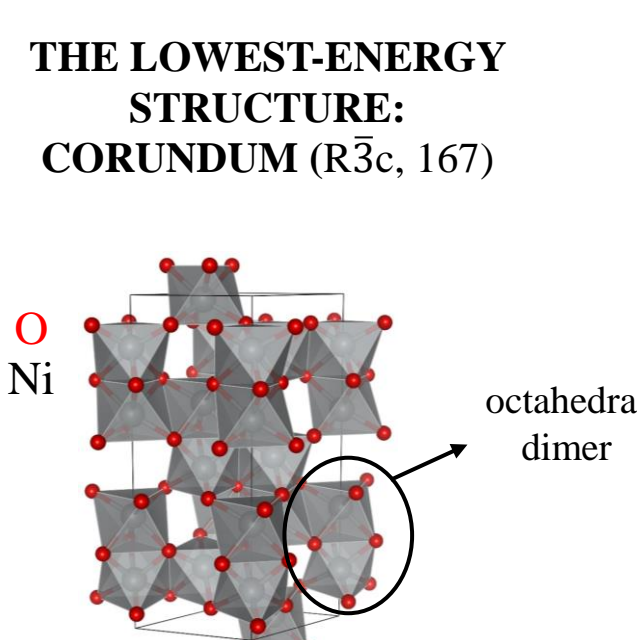
- This structure appears to be a ferromagnetic conductor

	Ni	O
C2/c	0.678 - 0.688	0.211 - 0.213, -0.085 - -0.211

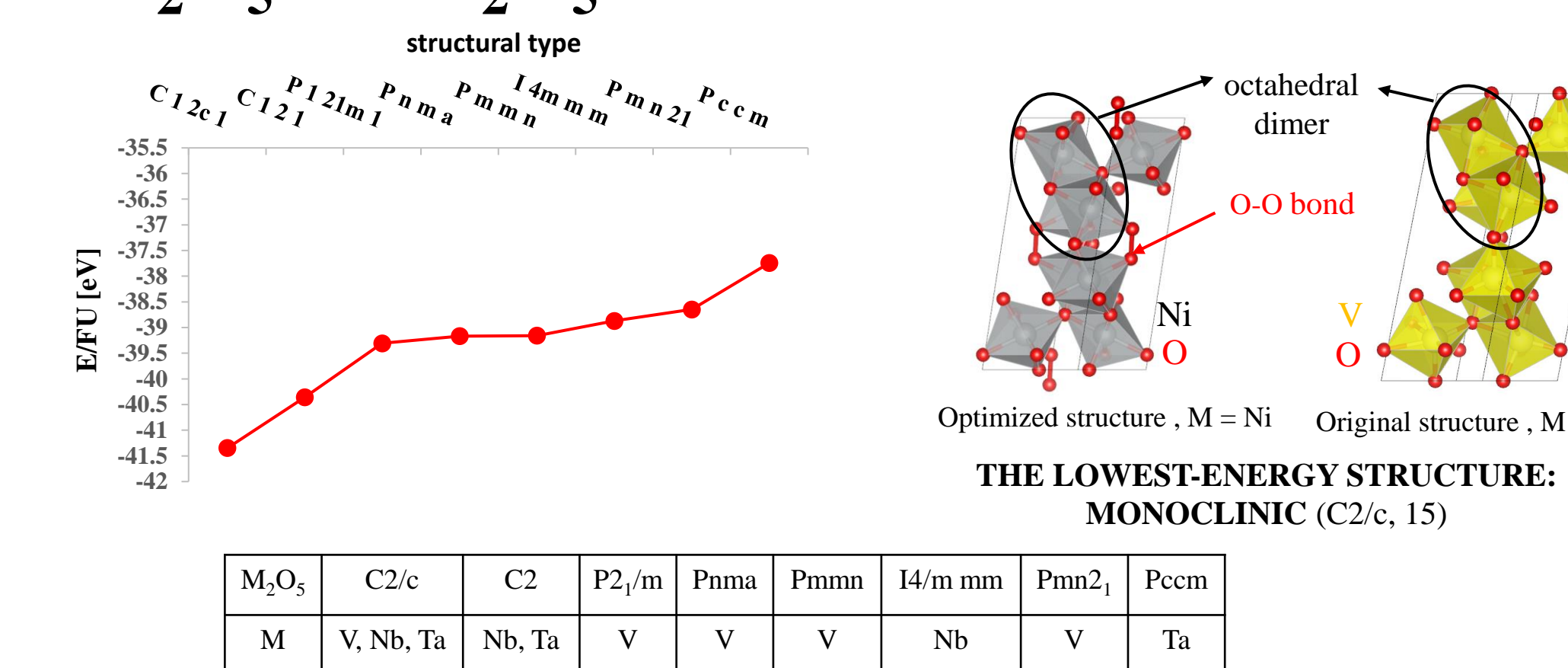
Ni₂O₃ IN M₂O₃ STRUCTURE TYPES



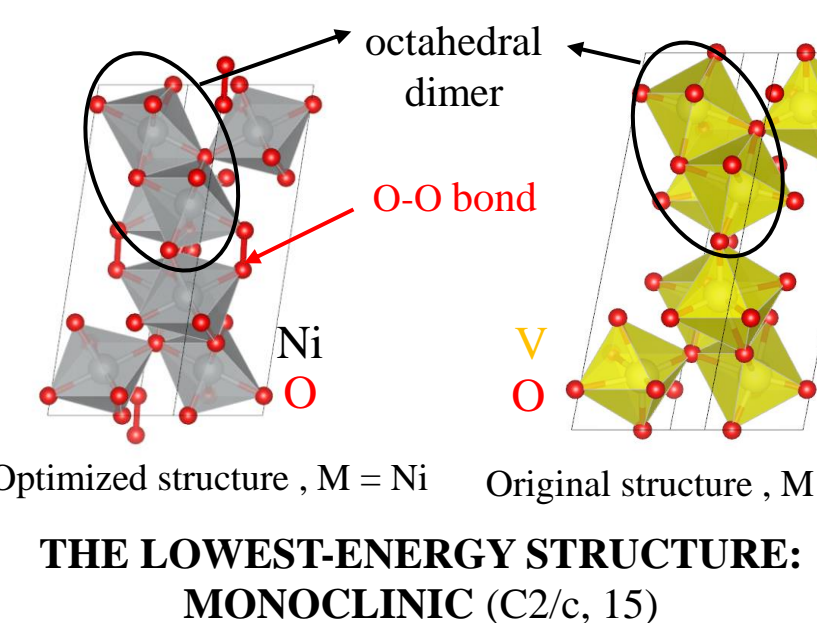
- Low-spin Ni³⁺ energetically preferred in all Ni₂O₃ models.
- The lowest energy structure is corundum type.
- Characteristic feature of corundum structure are **octahedral dimers** (in the c-direction) formed by MO₆ octahedrons that share a common face.



Ni₂O₅ IN M₂O₅ STRUCTURE TYPES

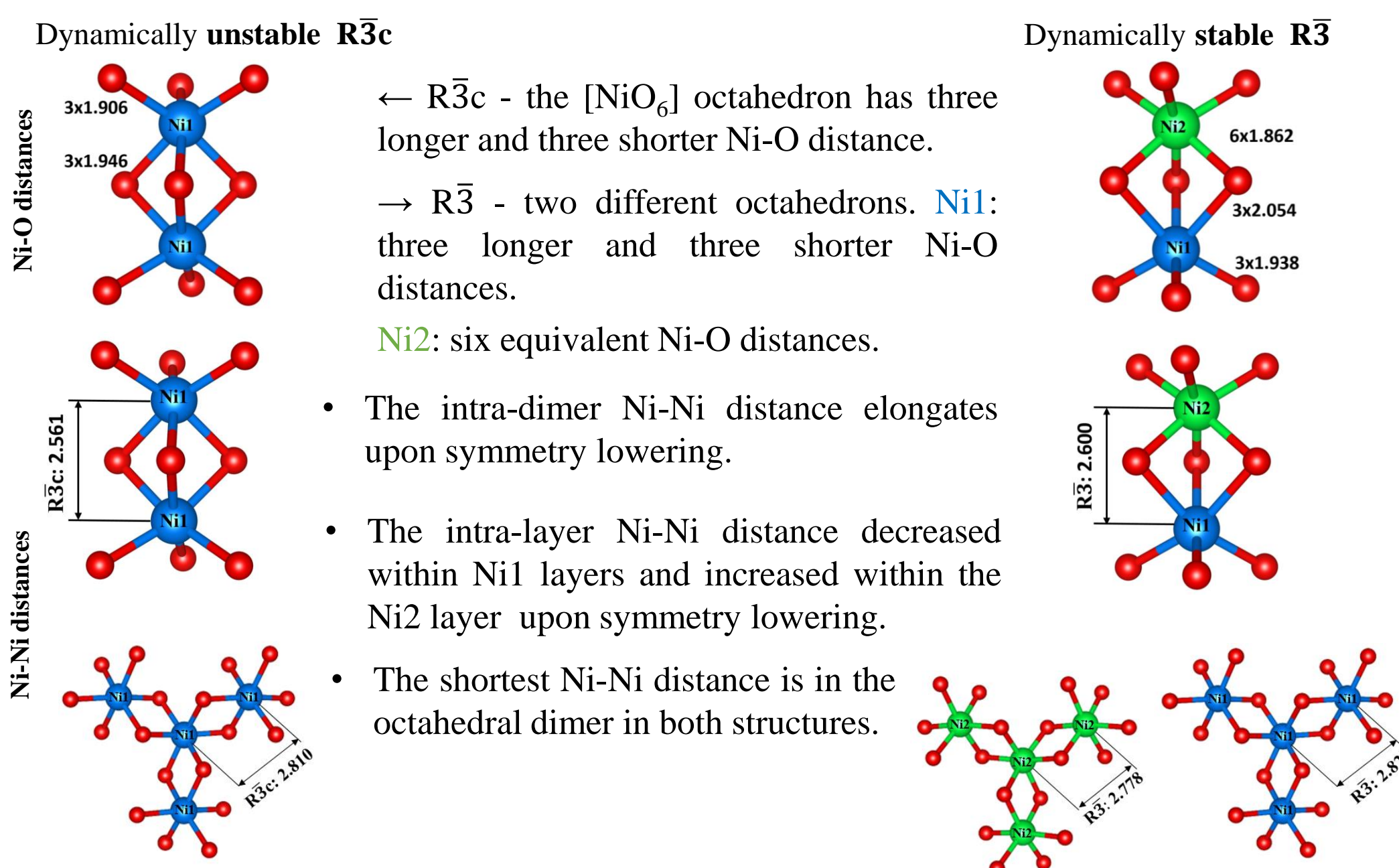


- The structure with lowest energy (C2/c, 15) is common for all known M₂O₅ oxides (see table above)
- [NiO₆] octahedra from **octahedral dimers** by sharing a common edge. The dimers are interconnected by oxygen atoms in shared vertices.
- O-O bonds are being formed in all optimized Ni₂O₅ modes in contrary to the original structures commonly taken by V, Nb and Ta pentoxides.



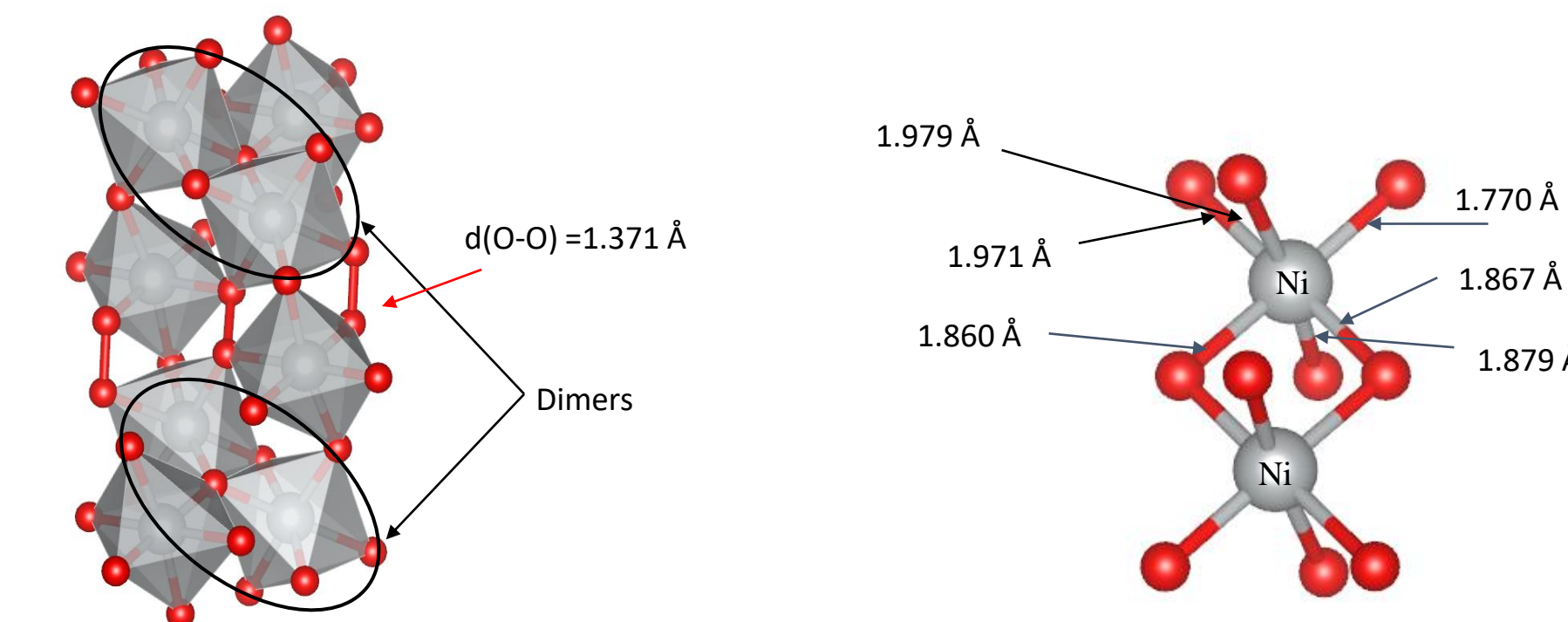
M ₂ O ₅	C2/c	C2	P2 ₁ /m	Pnma	Pmnm	I4/m mm	Pmm2 ₁	Pccm
M	V, Nb, Ta	Nb, Ta	V	V	V	Nb	V	Ta

ANALYSIS OF GEOMETRY



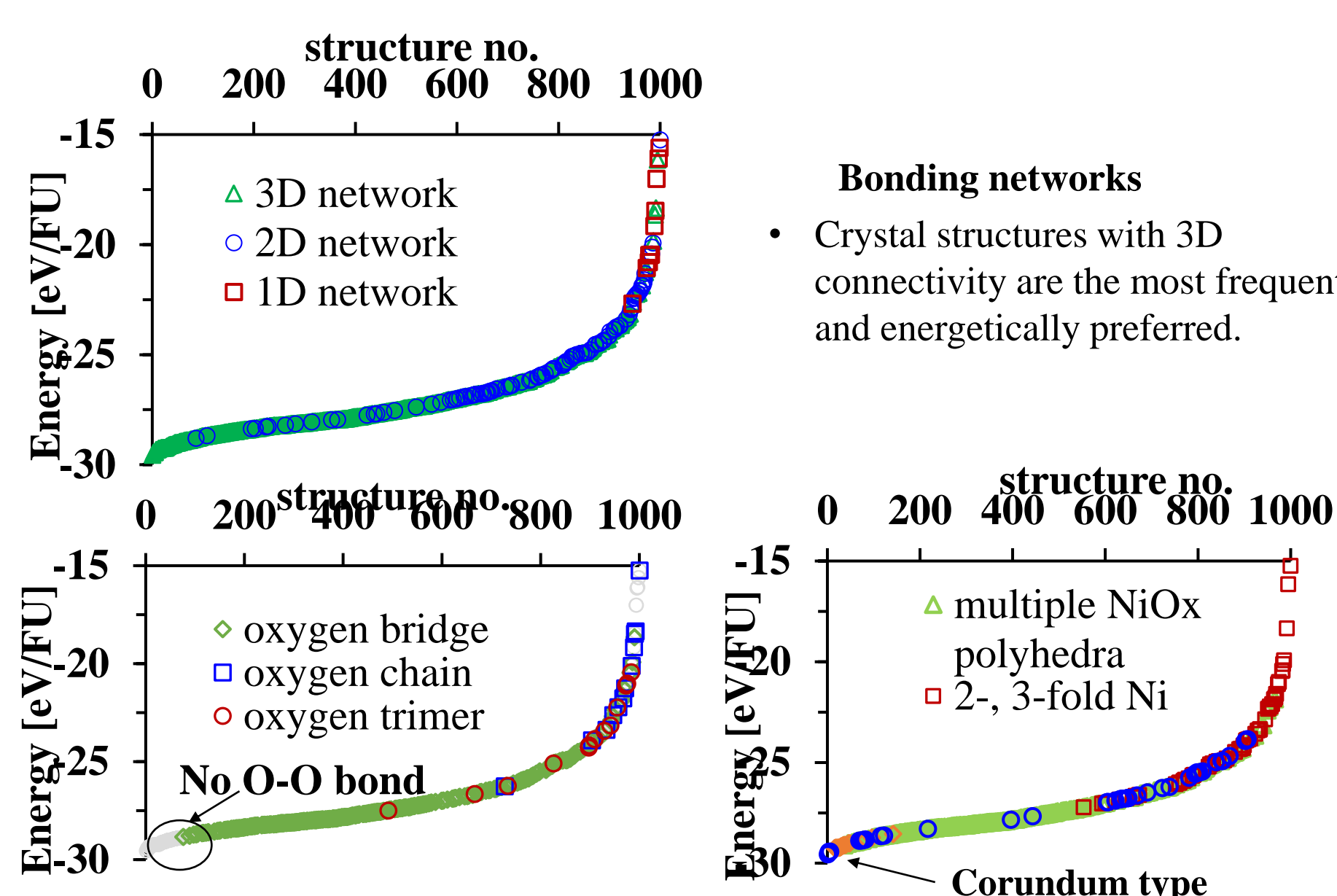
- The intra-dimer Ni-Ni distance elongates upon symmetry lowering.
- The intra-layer Ni-Ni distance decreased within Ni1 layers and increased within the Ni2 layer upon symmetry lowering.
- The shortest Ni-Ni distance is in the octahedral dimer in both structures.

ANALYSIS OF GEOMETRY



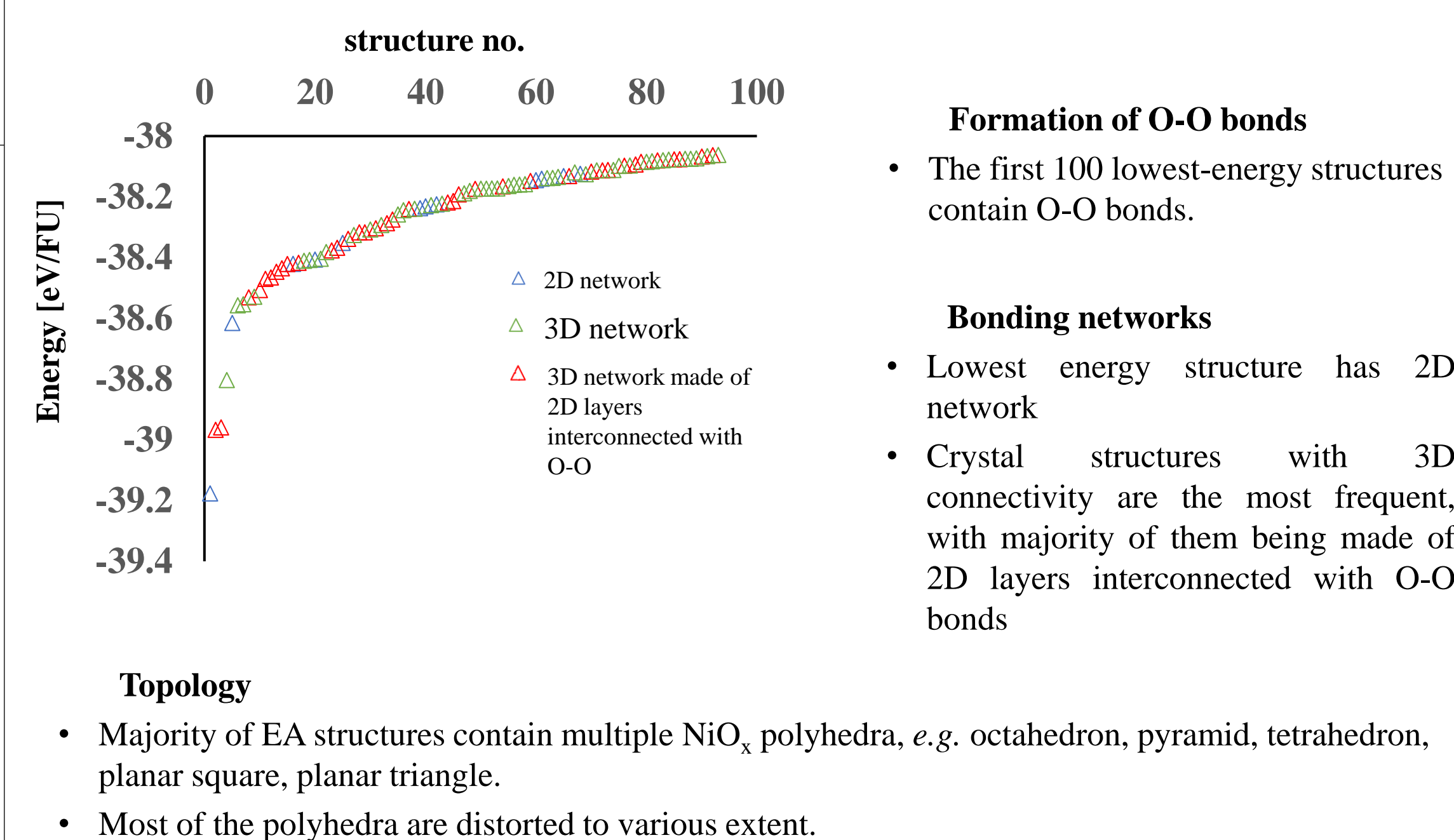
- The intra-dimer Ni-Ni distance is 2.833 Å.
- Dimers are made of slightly distorted octahedrons - all distances between Ni and coordinating O atoms in one octahedron are different.
- Very short O-O distances are present in this structure.
- Octahedron dimers share vertices and are further interconnected by O-O bridges.

EVOLUTIONARY ALGORITHMS



- Bonding networks**
 - Crystal structures with 3D connectivity are the most frequent and energetically preferred.
- Formation of O-O bonds**
 - Majority of structures contain oxygen-oxygen bonds.
 - The lowest-energy structures **do not** contain oxygen-oxygen bonds.
- Topology**
 - Majority of EA structures contain multiple NiO_x polyhedra: e.g. octahedron, pyramid, square, butterfly.
 - The lowest-E structure is **corundum**.

EVOLUTIONARY ALGORITHMS



- Formation of O-O bonds**
 - The first 100 lowest-energy structures contain O-O bonds.
- Bonding networks**
 - Lowest energy structure has 2D network
 - Crystal structures with 3D connectivity are the most frequent, with majority of them being made of 2D layers interconnected with O-O bonds
- Topology**
 - Majority of EA structures contain multiple NiO_x polyhedra, e.g. octahedron, pyramid, tetrahedron, planar square, planar triangle.
 - Most of the polyhedra are distorted to various extent.

REFERENCES

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